

PCT

WORLD INTELLECTUAL PROPERTY ORGANIZATION
International Bureau



INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification 7 : C10L 1/02, 1/14	A1	(11) International Publication Number: WO 00/36055 (43) International Publication Date: 22 June 2000 (22.06.00)
(21) International Application Number: PCT/GB99/04155 (22) International Filing Date: 15 December 1999 (15.12.99) (30) Priority Data: 9827592.8 15 December 1998 (15.12.98) GB (71) Applicant (for all designated States except US): AAE HOLDINGS PLC [GB/GB]; Unit 11, Bridge Road, Business Park, Haywards Heath, West Sussex RH16 1TX (GB). (72) Inventors; and (75) Inventors/Applicants (for US only): HAZEL, Clifford, James [GB/GB]; Keveral Mill, Hessenford, Cornwall PL11 3HW (GB). WILLIAMSON, Ian, Vernon [GB/GB]; 27 Pownall Avenue, Bramhall, Stockport, Cheshire SK7 2HE (GB). (74) Agent: GILHOLM, Steve; Harrison Goddard Foote, Belmont House, 20 Wood Lane, Leeds LS6 2AE (GB).		(81) Designated States: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZA, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SL, SZ, TZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG). Published <i>With international search report.</i> <i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i>
(54) Title: FUEL COMPOSITION (57) Abstract There is described a method of reducing the vapour pressure (RVP) of a gasoline/alcohol mixture which comprises adding a surfactant composition comprising an alkanolamide, an alkoxylated alcohol and an alkoxylated fatty acid to a gasoline/alcohol mixture.		

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AL	Albania	ES	Spain	LS	Lesotho	SI	Slovenia
AM	Armenia	FI	Finland	LT	Lithuania	SK	Slovakia
AT	Austria	FR	France	LU	Luxembourg	SN	Senegal
AU	Australia	GA	Gabon	LV	Latvia	SZ	Swaziland
AZ	Azerbaijan	GB	United Kingdom	MC	Monaco	TD	Chad
BA	Bosnia and Herzegovina	GE	Georgia	MD	Republic of Moldova	TG	Togo
BB	Barbados	GH	Ghana	MG	Madagascar	TJ	Tajikistan
BE	Belgium	GN	Guinea	MK	The former Yugoslav Republic of Macedonia	TM	Turkmenistan
BF	Burkina Faso	GR	Greece			TR	Turkey
BG	Bulgaria	HU	Hungary	ML	Mali	TT	Trinidad and Tobago
BJ	Benin	IE	Ireland	MN	Mongolia	UA	Ukraine
BR	Brazil	IL	Israel	MR	Mauritania	UG	Uganda
BY	Belarus	IS	Iceland	MW	Malawi	US	United States of America
CA	Canada	IT	Italy	MX	Mexico	UZ	Uzbekistan
CF	Central African Republic	JP	Japan	NE	Niger	VN	Viet Nam
CG	Congo	KE	Kenya	NL	Netherlands	YU	Yugoslavia
CH	Switzerland	KG	Kyrgyzstan	NO	Norway	ZW	Zimbabwe
CI	Côte d'Ivoire	KP	Democratic People's Republic of Korea	NZ	New Zealand		
CM	Cameroon			PL	Poland		
CN	China	KR	Republic of Korea	PT	Portugal		
CU	Cuba	KZ	Kazakhstan	RO	Romania		
CZ	Czech Republic	LC	Saint Lucia	RU	Russian Federation		
DE	Germany	LI	Liechtenstein	SD	Sudan		
DK	Denmark	LK	Sri Lanka	SE	Sweden		
EE	Estonia	LR	Liberia	SG	Singapore		

FUEL COMPOSITION

The invention relates to a novel method of reducing the vapour pressure of fuels and to the new use of compositions.

5

It is well known that liquid fuels when burned in an internal combustion engine can give rise to pollution and other undesired side effects. Furthermore, the relatively high vapour pressure of most liquid hydrocarbon fuels is known to give rise to environmental and safety problems upon storage and transportation. Numerous proposals have been advanced to reduce these side effects and enhance efficiency, eg miles per gallon. It has been realised that surfactants can play a useful role in this context but so far as we are aware none has satisfied the modern commercial criteria. It is one object of this invention to meet the need.

15 According to the US Environmental Protection Agency (EPA), volatile organic compound (VOC) emissions from gasoline are, as ozone precursors, a major contributor to the nation's serious ground level ozone problem, which harms human health and the public welfare. In addition to automotive emissions, gasoline evaporation occurs during refining, during commercial handling, transportation, and storage, and during refuelling. Because gasoline evaporation is a significant environmental problem, work was initiated at the National Alternative Fuels Laboratory (NAFL) of the University of North Dakota Energy & Environmental Research Center (EERC) to investigate evaporation rates of various gasolines and gasoline-oxygenate blends and the composition of their evaporative emissions. One aspect of the work, which was funded by the US Department of Agriculture (USDA), was to compare evaporation rates and compositions of E10 evaporative emissions with those of nonethanol-blended base fuels.

Widespread commercial use of E10 fuels was originally promoted as a means to reduce gasoline consumption during the oil embargo of 1973. Current interest in ethanol as a transportation fuel involves its use as an oxygenate to reduce carbon

monoxide (CO) combustion emissions. Also, its octane-boosting quality makes it useful as a replacement for octane-supplying aromatics that are slated for removal from reformulated gasolines to be sold in ozone non-attainment areas. Despite its desirable qualities as a fuel and the fact that it is renewable, ethanol has serious competition. Ethanol supporters are locked in a long-term debate with supporters of the other major commercially available gasoline oxygenate, methyl *tertiary*-butyl ether (MTBE), over which oxygenate is best for the environment.

A major point of contention in the debate is the approximate one-pound per square inch (psi) increase in RVP that occurs upon addition of ethanol to base fuel at a concentration of 10 vol%. However, while RVP is a useful indicator of a fuel's volatility, it is less useful as an indicator of how atmospheric interaction of vapour from the fuel (generated via evaporation or displacement during tank filling) may affect the environment in terms of ground level ozone production. Because gasolines are complex mixtures of over 200 components, the overall ozone-producing potential of a particular gasoline's vapour will depend not only on how much vapour is emitted, but also on the ozone-producing potential and concentration of each vapour component.

Surfactants have long had utility as additives which can affect the performance of fuels such as gasoline and diesel. British Patent No 2217229 describes a solubilising compound as a fuel additive. In particular there is described a composition comprising;

- 48 parts by volume of an alcohol ethoxylate;
- 3-8 parts by volume of lauric diethanolamide;
- 3-8 parts by volume of oleic diethanolamide; and
- 1.5-4 parts by volume of a polyglycol ether such as ethoxylated oleic acid.

Such compositions are useful as fuel additives and enable the solubilisation of water in fuels thus reducing its potential corrosive effect. However, the compositions are disadvantageous in that, *inter alia*, they require a high additive to fuel ratio.

Furthermore, they do not address the problems of emissions of gases such as CO, CO₂ and NOX, nor do they address the problems of evaporative loss due to the high vapour pressure of the fuel composition.

- 5 International Patent Application No WO 98/17745 describes an alternative formulation which comprises,

25% w/w of a diethanolamide,
50% w/w of an ethoxylated alcohol, and
10 25% w/w of a fourteen carbon chain fatty acid with seven ethoxylate groups.

WO '745 especially describes fuel compositions comprising, *inter alia*, an additive made up of a fatty acid diethanolamide, an alcohol ethoxylate and an ethoxylate of a fatty acid, the degree of ethoxylation being selected so that a long term stable fuel
15 composition is formed and, in particular, wherein, by carefully selecting the degree of ethoxylation, a balanced blend can be produced, such that the molecular weight ratios of each of the three components are substantially equivalent, for example, the molecular weight ratios are 1:1:1.

- 20 Whilst such additives provide significant reductions in emissions and are useable at low concentrations, there is still a need for a fuel composition which is capable of reducing emissions whilst maintaining performance levels.

It is known that to meet some legislation requirements fuels must have a low Reid
25 Vapour Pressure value (RVP), e.g. under California State law, the RVP must be below 7psi.

Some modern day gasoline contains a small percentage of a condensate from natural gas pipelines. Gasoline is made up from a range of hydrocarbons, which is
30 determined by the refining and blending processes. They may also contain minor amounts of comparatively volatile contaminants picked up during transportation, for

example, through pipe lines which have been previously used for transportation of natural gases. Pentanes are a range of C5 hydrocarbons that are typically left in the crude oil cut after the gas condensates are removed. Natural gas (methane) is extracted first from the oil field, ethane is increasingly also extracted at the well for petrochemical production of ethylene, while LPG, a mixture of C3 and C4 hydrocarbons, is liquefied under moderate pressure for use as a gas fuel. Pentane is usually the first cut off the Crude Distillation Unit at the refinery and the disposal of pentane creates a problem for the petroleum industry. Pentane has little value as a petrochemical feedstock, and blending it into the gasoline pool has the undesirable effect of raising the RVP. The aforementioned contaminants, as well as low boiling fractions arising from the refining and blending processes, all have the undesirable effect of increasing the vapour pressure of the gasoline. This undesirable effect will be further exacerbated if the fuel is blended with certain volatile oxygenating agents such as ethanol. Thus the condensate is a light fraction which is primarily pentanes although other hydrocarbons may be present. The condensate adversely affects the RVP, especially when the gasoline contains an alcohol such as ethanol.

It is known that when gasoline and an alcohol are mixed, the resultant mixture has an increased vapour pressure. We have now surprisingly found that by adding certain surfactant compositions to a gasoline and alcohol mixture a reduction in vapour pressure (RVP) can be achieved.

Thus, according to the invention we provide a method of reducing the vapour pressure of a gasoline/alcohol mixture which comprises adding a surfactant composition comprising an alkanolamide, an alkoxyated alcohol and an alkoxyated fatty acid to a gasoline/alcohol mixture.

By the term vapour pressure we mean the Reid Vapour Pressure (RVP). Thus the method of the invention preferably reduces the RVP of the gasoline mixture to less than 8psi, preferably less than 7psi and especially between 6 and 7psi.

For the purposes of this invention the RVP is measured using the test method of ASTM D5191. Thus the vapour pressure is given as psi at 37.8°C.

It is a novel aspect of the invention to use the aforementioned surfactant composition in the reduction of RVP.

According to a further aspect of the invention we provide the use of an additive composition as hereinbefore described in the manufacture of a fuel composition having an RVP of less than 8psi.

10

The alkanolamide is preferably an ethanolamide and more preferably a diethanolamide. Especially preferred are the diethanolamides and particularly the super diethanolamides. By the term super diethanolamide we mean a diethanolamide in which the nitrogen is substituted by an acyl substituent e.g. acyl C₅ to C₂₀, preferably C₈ to C₁₆, more preferably C₁₀ to C₁₅. The most preferred diethanolamide has a C₁₂ acyl substituent i.e. lauryl diethanolamide. Where the amides are derived from natural products this moiety will have an even number of carbon atoms, e.g. 12 for the lauryl derivative. Note, the alkyl part of this group is the R group which will be an odd number.

20

There are three commercial routes to alkanolamides;

Acid + alkanolamine = alkanolamide + water

Plant or animal oil (triglyceride) + alkanolamine = alkanolamide + glycerol

25

Methyl ester + alkanolamine = alkanolamide + methanol

These are listed in order of increasing product quality. The route via the acid often uses an excess of alkanolamine to produce a product higher in amide than is obtainable from the acid if a stoichiometric ratio is used; these products are sometimes referred to as super amides.

30

The alkoxyated alcohol is preferably an ethoxylated alcohol. It is essential that the ethoxylated alcohol is an oil soluble alcohol. Therefore, alkanols are preferred and these may be primary, secondary or tertiary alkanols and especially primary alkanols. As the oil solubility of the alcohol may vary with the carbon chain length of the ethoxylated alkanol, the alkanol is preferably a C₅ to C₂₂ alkanol, more preferably C₅ to C₁₅ alkanol. The ethoxylated alcohol may comprise a mixture of alkanols. However, it is preferred that in such mixtures one alkanol will predominate. Thus, the most preferred alkanol is predominantly a C₉ to C₁₁ alkanol. In addition the degree of ethoxylation of the alcohol may be varied and the oil solubility will, generally, decrease with the increase in the degree of ethoxylation. It is preferred that the ethoxylate to alcohol ratio is greater than 2. More preferably, the ethoxylate to alcohol ratio is from between 1 and 10, preferably between 1 and 5, more preferably between 1 and 3 and especially between 2 and 3. A commercially available ethoxylated alcohol is especially preferred in which the ethoxylate to alcohol ratio is 2.75. Such an alcohol ethoxylate is available as NEODOL 91/2.5.

The fatty acid ethoxylate may comprise any conventionally known fatty acid ethoxylate. Thus the fatty acid ethoxylate may be derived from a fatty acid having from 8 to 20 carbon atoms, preferably from 10 to 18 and most preferably 14 carbon atoms (myristic acid).

The degree of ethoxylation is chosen to optimise performance in the blend with the other two selected surfactants and may be from 1 to 20, but more preferably from 5 to 12. A suitable product within this range would be, for example that derived from the addition of 7 molecules of ethylene oxide too 1 mole of myristic acid.

The values will depend on the length of the hydrophilic chain, typically an ethoxylate chain. The length of the chain will increase the extent of solubilisation because of a greater ability to solubilise.

As with the compositions described in WO98/17745, a blend of surfactants is preferred, preferably by selecting one appropriate to the fuel, say 10 to 18 for hydrocarbon fuel, most preferably 13. In the case of an alcohol the HLB value of the surfactant is between 3 and 7, most preferably about 4. But the addition of
5 surfactants normally create ratios of 1:1 or high volume emulsions or 5:1 ratios when the solubalisation is required at 1:100.

The invention has the ability to unify the HLB requirements of any liquid fuel which in turn allows for one dose to be used in any fuel from C5 carbon chains up. The
10 benefit being the amount of treatment directly related to the co-solvency ability.

Preferably the ethoxylate of the fatty acid makes up about 25% by volume of the additive and further preferably the alcohol ethoxylate comprises 50% by volume of the additive.
15

The surfactant additive may be added to a hydrocarbon fuel, eg diesel, petrol or alcohol, such as ethanol which may or may not be contaminated with water.

The fuel composition comprising a fuel and a fuel additive wherein the additive to
20 fuel ratio is 0.5 – 1:1200.

The alcohol to hydrocarbon fuel ratio is from 1:8 to 1:9. Water may be present, e.g. as condensate, or added separately, in water: alcohol: fuel ratio of 0:1:8 to 0.5:1:9 and preferably from 0.25:1:8 to 0.25:1:9.
25

The additive may be present in an additive to fuel/alcohol ratio of about 0.5 – 1:1000. Preferably the additive to fuel ratio is about 1:1000, most preferably about 1:1200. In this context the fuel is the hydrocarbon fuel and alcohol mixture.

Example 1

Different blends shown in Table 1 were made and subjected to gas chromatography. The ratio of hydrocarbon: ethanol was measured by gas chromatography. The results were used to indicate the extent to which the content of the ethanol could be incorporated without exceeding an increase in the temperature at which the volatile ingredients evaporated. The results showed the amount of ethanol that could be incorporated into the hydrocarbon without increasing the temperature at which the volatile components evaporated.

The additive comprised 60% by weight ethanolethoxylate, 20% by weight diethanolamide and 20% by weight of lauric diethanolamide. Our investigation suggests that these interact to form a blend in which the ethanolamide is a backbone.

Example 2

Different blends were made up and the RVP determined. The following results were obtained:

Table II

Sample	Gasoline	Ethanol	Additive as above	Water	RVP (1)	RVP (2)	RVP (3)
1	90	10	-	-	7.48	7.7	7.58
2	87.5	10	2.5	-	7.18	7.41	7.29
3	85.0	10	2.5	2.5	6.87	7.11	6.98
4	85.0	10	2.5	2.5	6.96	7.19	7.06

The RVP value was measured according to ASTM D5191 and is the mean value of results calculated according to CARB (1), EPA (2) and ASTM (3) methods. In Samples 2, 3 and 4 the gasoline included a proportion of a condensate which was mainly pentanes. These results demonstrate a tendency of the additive to reduce the

RVP value, and that by selecting appropriate proportions a fuel composition can be made to meet the requirements of local legislation.

Table 1

Sample	Hydrocarbon (a)	Ethanol (b)	Pentane	Additive	Ratio (a):(b)
1	90	10	0	0	15:1
2	89	10	0	1	16.7:1
3	82	13.5	4.5	0	17.75:1
4	77.5	13.5	4.5	4.5	11.69:1

5

Example 3

Various tests were conducted according to ASTM D-5191. The results are given in
10 Tables III to IX.

Example 4

15 Various tests were conducted according to the 1975 US Federal City Gasoline Test.
The results are given in Tables X to XIII;

Table X: Tests on CARB Gasoline

Table XI: Tests on EPA

20 Table XII: Tests on Gasohol

25

CLAIMS

1. A method of reducing the vapour pressure of a gasoline/alcohol mixture which comprises adding surfactant composition comprising an alkanolamide, an
5 alkoxyated alcohol and an alkoxyated fatty acid to a gasoline/alcohol mixture.
2. A method according to Claim 1 characterised in that the RVP is less than 7psi.
- 10 3. A method according to Claim 2 characterised in that the RVP is less than 7psi.
4. A surfactant composition according to Claim 3 characterised in that the RVP is between 6 and 7psi.
- 15 5. A method according to Claim 1 characterised in that the alkanolamide is a diethanolamide.
6. A method according to Claim 5 characterised in that the nitrogen in the
20 diethanolamide is substituted by an alkyl C₅ to C₂₀ substituent.
7. A surfactant composition according to Claim 6 characterised in that the diethanolamide is a lauryl diethanolamide.
- 25 8. A method according to Claim 1 characterised in that the alkoxyated alcohol is an ethoxyated alcohol.
9. A method according to Claim 8 characterised in that the ethoxyated alcohol is a C₅ to C₁₅ alkanol.

30

10. A method according to Claim 8 characterised in that the ethoxylated alcohol comprises a mixture of alkanols in which one alkanol predominates.
11. A method according to Claim 8 characterised in that the predominate alkanol
5 is a C₉ to C₁₁ alkanol.
12. A method according to Claim 8 characterised in that the ethoxylate to alcohol ratio is from between 1 and 5.
- 10 13. A method according to Claim 8 characterised in that the ethoxylated alcohol is NEODOL 91/2.5.
14. A method according to Claim 1 characterised in that the fatty acid group is a C₈ to C₂₀ fatty acid.
15
15. A method according to Claim 14 characterised in that the fatty acid group is a C₁₄ fatty acid (myristic acid).
16. A method according to Claim 1 characterised in that the ester moiety of the
20 fatty acid ester is an alkyl ester.
17. A method according to Claim 16 characterised in that the alkyl group is a C₁ to C₁₀ alkyl.
- 25 18. A method according to Claim 1 characterised in that the composition comprises 25% v/v of the fatty acid ester.
19. A method according to Claim 1 characterised in that the composition comprises 50% v/v of the alcohol ethoxylate.
30

20. A method according to claim 1 characterised in that the surfactant additive to fuel/alcohol ratio is from 0.5:1200 to 1:1000.

21. The use of a surfactant composition comprising an alkanolamide, an alkoxyated alcohol and an alkoxyated fatty acid ester in the manufacture of a fuel composition having an RVP of less than 8psi.

22. A method of manufacturing a mixture comprising gasoline, alcohol and a surfactant composition, said surfactant composition comprising an alkanolamide, an ethoxyated alcohol and an alkoxyated fatty acid characterised in that the method comprises blending the alcohol and surfactant followed by blending with gasoline.

23. A method substantially as described with reference to the accompanying examples.

1/10

TABLE III

LABORATORY TESTS RESULTS

LABORATORY TESTS RESULTS		05/28/99			
JOB NUMBER: 990554		ATTN: Attn			
CLIENT I.D.: C - BASE CARBOB (GASOLINE)		LABORATORY I.D.: 990554-0034			
DATE SAMPLED: 05/26/99		DATE RECEIVED: 05/26/99			
TIME SAMPLED: 12:02		TIME RECEIVED: 12:02			
WORK DESCRIPTION: C		REMARKS:			
TEST DESCRIPTION	FINAL RESULT	LIMITS* DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/26/99
RVP CARB CALCULATION	5.82*	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	6.08	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	5.94	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	6.73	0.10	PSI	ASTM D-5191	

2/10

TABLE IV

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn		
CLIENT I.D.....	CARBOB 90/10	BASE + 10% ETOH	LABORATORY I.D.....	990554-0039
DATE SAMPLED.....	05/20/99		DATE RECEIVED.....	05/20/99
TIME SAMPLED.....	16:01		TIME RECEIVED.....	16:01
WORK DESCRIPTION....	CARBOB 90/10		REMARKS.....	1-500ml amber btl
TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191
RVP CARB CALCULATION	7.16*	0.10	PSI	ASTM D-5191 (CARB)
RVP EPA CALCULATION	7.41	0.10	PSI	ASTM D-5191 (EPA)
RVP ASTM CALCULATION	7.29	0.10	PSI	ASTM D-5191 (ASTM)
PRESSURE TOTAL/GRABNER INSTRUMENT	8.12	0.10	PSI	ASTM D-5191
				05/28/99

BEST AVAILABLE COPY

3/10

TABLE V

LABORATORY TESTS RESULTS

05/28/99

JOB NUMBER: 990554

ATTN: Attn

CLIENT I.D.: CARBOB 90/10 LN
 DATE SAMPLED: 05/20/99
 TIME SAMPLED: 16:01
 WORK DESCRIPTION: CARBOB 90/10 LN
 BASE + 10% ETOH
 + 1% 25% NEODOL
 37.5% OPE
 37.5% ATLAS
 LABORATORY I.D.: 990554-0040
 DATE RECEIVED: 05/20/99
 TIME RECEIVED: 16:01
 REMARKS: 1-500ml amber btl

TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.79	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.04	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.91	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.73	0.10	PSI	ASTM D-5191	

BEST AVAILABLE COPY

4/10

TABLE VI

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn		
CLIENT I.D.....	CARBOB 90/10 REVERSE	BASE + 10% ETOH	LABORATORY I.D.....	990554-0042
DATE SAMPLED.....	05/20/99	+ 1% AAE + 75% NEODOL	DATE RECEIVED.....	05/20/99
TIME SAMPLED.....	16:01	+ 12.5% ODE	TIME RECEIVED.....	16:01
WORK DESCRIPTION....	CARBOB 90/10 REVERSE	+ 12.5% ATLAS	REMARKS.....	1-500ml amber btl
TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191
RVP CARB CALCULATION	6.56	0.10	PSI	ASTM D-5191 (CARB)
RVP EPA CALCULATION	6.80	0.10	PSI	ASTM D-5191 (EPA)
RVP ASTM CALCULATION	6.67	0.10	PSI	ASTM D-5191 (ASTM)
PRESSURE TOTAL/GRABNER INSTRUMENT	7.48	0.10	PSI	ASTM D-5191
				05/28/99

5/10

TABLE VII

LABORATORY TESTS RESULTS

05/28/99

JOB NUMBER: 990554

ATTN: Attn

CLIENT I.D.: CARBOB 90/10 (RECHECK)

LABORATORY I.D.: 990554-0043

DATE SAMPLED: 05/27/99

DATE RECEIVED: 05/27/99

TIME SAMPLED: 16:01

TIME RECEIVED: 16:01

WORK DESCRIPTION: CARBOB 90/10 (RECHECK)

REMARKS: 1-500ml amber btl

TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	7.22	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.47	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	7.34	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	8.18	0.10	PSI	ASTM D-5191	

6/10

TABLE VIII

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn		
CLIENT I.D.....	CARBOB 90/10 LN (RECHECK)	LABORATORY I.D.....	990554-0044	
DATE SAMPLED.....	05/27/99	DATE RECEIVED.....	05/27/99	
TIME SAMPLED.....	16:01	TIME RECEIVED.....	16:01	
WORK DESCRIPTION....	CARBOB 90/10 LN (RECHECK)	REMARKS.....	1-500ml amber btl	

TEST DESCRIPTION	FINAL RESULT	UNITS/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.76	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	7.01	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.88	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.70	0.10	PSI	ASTM D-5191	

7/10

TABLE IX

LABORATORY TESTS RESULTS				
05/28/99				
JOB NUMBER: 990554		ATTN: Attn.		
CLIENT I.D.: CARBOB 90/10 REVERSE (RECHECK)		LABORATORY I.D.: 990554-0046		
DATE SAMPLED: 05/27/99		DATE RECEIVED: 05/27/99		
TIME SAMPLED: 16:01		TIME RECEIVED: 16:01		
WORK DESCRIPTION: CARBOB 90/10 REVERSE (RECHECK)		REMARKS: 1-500ml amber btl		

TEST DESCRIPTION	FINAL RESULT	LIMITS/DILUTION	UNITS OF MEASURE	TEST METHOD	DATE
REID EQUIVALENT VAPOUR PRESSURE		*1		ASTM D-5191	05/28/99
RVP CARB CALCULATION	6.56	0.10	PSI	ASTM D-5191 (CARB)	
RVP EPA CALCULATION	6.80	0.10	PSI	ASTM D-5191 (EPA)	
RVP ASTM CALCULATION	6.67	0.10	PSI	ASTM D-5191 (ASTM)	
PRESSURE TOTAL/GRABNER INSTRUMENT	7.48	0.10	PSI	ASTM D-5191	

8/10

TABLE X

Fuel Summary CARB Gasoline vs CARB Gasoline + AAE10 @ 1000ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	Nox g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5736	1993 Ford Explorer	CARB Base, Seq#1	0.151	1.697	0.239	524.16	0.027	0.1196	16.183
5736	1996 GMC Jimmy	CARB Base, Seq#1	0.218	2.744	0.288	485.79	0.03	0.1835	17.387
5675	1998 Dodge Ram	CARB Base, Seq#1	0.135	1.427	0.132	582.37	0.021	0.1108	14.586
5682	1991 Geo Prizm	CARB Base, Seq#1	0.348	3.269	0.372	289.68	0.031	0.3129	28.84
5736	1996 Nissan Sentra	CARB Base, Seq#1	0.073	0.772	0.148	289.58	0.008	0.0638	29.323
5734	1994 Nissan Sentra	CARB Base, Seq#1	0.17	1.878	0.368	308.26	0.017	0.1511	27.373
		Average	0.183	1.964	0.258	413.305	0.022	0.157	22.282
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5691	1993 Ford Explorer	CARB + AAE10, Seq#2	0.18	1.586	0.245	532.46	0.029	0.1479	15.935
5684	1996 GMC Jimmy	CARB + AAE10, Seq#2	0.202	2.524	0.287	493.70	0.028	0.1695	17.125
5686	1998 Dodge Ram	CARB + AAE10, Seq#2	0.194	1.712	0.115	592.59	0.024	0.1662	14.32
5700	1991 Geo Prizm	CARB + AAE10, Seq#2	0.326	3.416	0.41	287.52	0.036	0.2846	29.035
5745	1996 Nissan Sentra	CARB + AAE10, Seq#2	0.074	0.785	0.129	294.62	0.008	0.0659	28.82
5747	1994 Nissan Sentra	CARB + AAE10, Seq#2	0.182	1.869	0.392	308.29	0.018	0.1613	27.369
		Average	0.193	1.982	0.263	418.199	0.024	0.166	22.101
		Deviation +/-	5.66%	0.89%	1.92%	1.18%	7.51%	5.70%	-0.81%

9/10

TABLE XI

Fuel Summary EPA vs EPA + AAE10 @ 1000ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5706	1993 Ford Explorer	EPA Base, Seq#3	0.223	2.316	0.34	539.82	0.043	0.1743	15.882
5693	1996 GMC Jimmy	EPA Base, Seq#3	0.242	2.882	0.453	499.87	0.042	0.1947	16.893
5692	1998 Dodge Ram	EPA Base, Seq#3	0.249	2.466	0.19	600.43	0.046	0.1965	14.103
5708	1991 Geo Prizm	EPA Base, Seq#3	0.327	3.392	0.446	294.90	0.035	0.2878	28.328
5752	1996 Nissan Sentra	EPA Base, Seq#3	0.165	1.038	0.208	299.19	0.012	0.1514	28.318
5753	1994 Nissan Sentra	EPA Base, Seq#3	0.253	2.952	0.333	315.84	0.023	0.2271	26.562
		Average	0.243	2.508	0.328	425.009	0.034	0.205	21.648
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5709	1993 Ford Explorer	EPA + AAE10, Seq#4	0.234	2.744	0.351	534.78	0.048	0.1804	15.808
5701	1996 GMC Jimmy	EPA + AAE10, Seq#4	0.324	3.141	0.517	500.54	0.046	0.2711	16.848
5702	1998 Dodge Ram	EPA + AAE10, Seq#4	0.315	3.308	0.227	598.06	0.056	0.2521	14.123
5717	1991 Geo Prizm	EPA + AAE10, Seq#4	0.365	3.952	0.519	294.89	0.034	0.3268	28.235
5758	1996 Nissan Sentra	EPA + AAE10, Seq#4	0.185	1.299	0.187	300.57	0.014	0.1697	28.146
5759	1994 Nissan Sentra	EPA + AAE10, Seq#4	0.30	3.308	0.371	313.13	0.025	0.2712	26.729
		Average	0.287	2.959	0.362	423.663	0.037	0.245	21.648
		Deviation +/-	18.00%	17.99%	10.27%	-0.32%	10.16%	19.45%	0.00%

10/10

TABLE XII

Fuel Summary Gasohol vs Gasohol + AAE00 @ 4600ppm

Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5722	1993 Ford Explorer	Gasohol Base, Seq#5	0.165	1.545	0.251	519.43	0.033	0.1275	16.336
5711	1996 GMC Jimmy	Gasohol Base, Seq#5	0.219	2.359	0.342	495.15	0.037	0.1765	17.083
5710	1998 Dodge Ram	Gasohol Base, Seq#5	0.219	2.102	0.194	589.97	0.047	0.1655	14.367
5727	1991 Geo Prizm	Gasohol Base, Seq#5	0.273	3.804	0.452	283.54	0.037	0.2309	29.549
5764	1996 Nissan Sentra	Gasohol Base, Seq#5	0.118	0.907	0.175	295.48	0.012	0.1046	28.706
5766	1994 Nissan Sentra	Gasohol Base, Seq#5	0.194	2.237	0.388	311.28	0.022	0.1694	27.056
		Average	0.198	1.992	0.30	415.807	0.032	0.162	22.183
Test No.	Vehicle	Fuel Tested EPA-75	THC g/mile	CO g/mile	NOx g/mile	CO2 g/mile	CH4 g/mile	NMHC g/mile	MPG
5726	1993 Ford Explorer	Gasohol + AAE00, Seq#6	0.17	1.48	0.26	515.20	0.035	0.1312	16.472
5770	1996 GMC Jimmy	Gasohol + AAE00, Seq#6	0.231	1.972	0.30	488.00	0.033	0.1938	17.351
5718	1998 Dodge Ram	Gasohol + AAE00, Seq#6	0.223	2.04	0.156	590.74	0.044	0.1733	14.351
5735	1991 Geo Prizm	Gasohol + AAE00, Seq#6	0.283	2.478	0.443	280.95	0.035	0.2436	29.866
5770	1996 Nissan Sentra	Gasohol + AAE00, Seq#6	0.096	0.793	0.135	294.61	0.01	0.0848	28.814
5769	1994 Nissan Sentra	Gasohol + AAE00, Seq#6	0.17	1.846	0.381	314.49	0.019	0.1492	26.842
		Average	0.196	1.768	0.279	413.998	0.029	0.163	22.282
		Deviation +/-	-1.25%	-11.25%	-7.24%	-0.44%	-7.22%	0.16%	0.45%

INTERNATIONAL SEARCH REPORT

International Application No

PCT/GB 99/04155

A. CLASSIFICATION OF SUBJECT MATTER
IPC 7 C10L1/02 C10L1/14

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C10L

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	WO 98 17745 A (WILLIAMSON IAN VERNON ;HAZEL CLIFFORD JAMES (GB)) 30 April 1998 (1998-04-30) cited in the application page 2, paragraph 2; claims 1,52; examples 2,6,10 page 4	1,5-15, 18-21,23
A	EP 0 466 511 A (ETHYL PETROLEUM ADDITIVES INC) 15 January 1992 (1992-01-15) page 5 -/--	21

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

* Special categories of cited documents :

"A" document defining the general state of the art which is not considered to be of particular relevance

"E" earlier document but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.

"&" document member of the same patent family

Date of the actual completion of the international search

19 April 2000

Date of mailing of the international search report

04/05/2000

Name and mailing address of the ISA

European Patent Office, P.B. 5818 Patentlaan 2
NL - 2280 HV Rijswijk
Tel. (+31-70) 340-2040, Tx. 31 651 epo nl,
Fax: (+31-70) 340-3016

Authorized officer

De La Morinerie, B

INTERNATIONAL SEARCH REPORT

International Application No

PCT/GB 99/04155

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	<p>DATABASE WPI Section Ch, Week 199641 Derwent Publications Ltd., London, GB; Class D25, AN 1996-408376 XP002136066 & JP 08 198830 A (KAWAKEN FINE CHEM CO LTD), 6 August 1996 (1996-08-06) abstract</p> <p>-----</p>	16,17

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/GB 99/04155

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
WO 9817745 A	30-04-1998	AU 4710097 A DE 19782068 T GB 2334964 A SE 9901364 A	15-05-1998 11-11-1999 08-09-1999 16-04-1999
EP 0466511 A	15-01-1992	AU 648564 B AU 8016791 A CA 2045455 A DE 69106611 D DE 69106611 T ES 2066357 T JP 4226598 A	28-04-1994 16-01-1992 14-01-1992 23-02-1995 18-05-1995 01-03-1995 17-08-1992
JP 8198830 A	06-08-1996	JP 2972395 B	08-11-1999

THIS PAGE BLANK (USPTO)